Table of Ions

Silver – Ag+

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Cations** | | | | | | | | |
|  | | 1+ | 2+ | | 3+ | | 4+ | |
| Stock System | | Copper I (Cu)  Hydronium (H3O) | Copper II (Cu)  Iron II (Fe)  Cobalt II (Co)  Chromium II (Cr)  Tin II (Sn)  Lead II (Pb) | | Iron III (Fe)  Cobalt III (Co)  Chromium III (Cr) | | Tin IV (Sn)  Lead IV (Pb) | |
| Polyatomic | | Ammonium (NH4)  Hydronium (H3O) |  | | | | | |
| **Anions** | | | | | | | | |
|  | 1- | | | 2- | | 3- | |  |
| Polyatomic | Acetate (C2H3O2)  Perchlorate (ClO4)  Chlorate (ClO3)  Chlorite (ClO2)  Cyanide (CN)  Dihydrogen phosphate (H2PO4)  Hydrogen carbonate (HCO3)  (aka bicarbonate)  Hydrogen sulfate (HSO4)  (aka bisulfate)  Hydroxide (OH)  Nitrate (NO3)  Nitrite (NO2)  Permanganate (MnO4) | | | Carbonate (CO3)  Chromate (CrO4)  Dichromate (Cr2O7)  Hydrogen phosphate (HPO4)  Oxalate (C2O4)  Peroxide (O2)  Sulfate (SO4)  Sulfite (SO3) | | Phosphate (PO4)  Phosphite (PO3) | |  |

You will need to memorize these ions by Tuesday 1/6 for the extra-credit quiz

How to read this table:

The element or group of elements in parenthesis carries the charge indicated in the column heading.

Examples: Tin IV Sn4+ Oxalate C2O42-